

ENGG1811 Computing for Engineers

Week 7 (Additional Slides)

For “Example – rates of diffusion”

Example – rates of diffusion*

- Metals are hardened by *carburising*, where carbon diffuses into the heated metal at a rate D that depends on the temperature T (in K) and material characteristics:

$$D = D_0 e^{-Q/RT}$$

- R is the ideal gas constant = 8.314 J/K/mol
- diffusion coefficient D_0 and activation energy Q depend on the material

Material	D_0 (m ² /s)	Q (J/mol)
Ferrite (α Fe)	6.2×10^{-7}	80 000

* Moore, Example 5.3, using SI units and including some corrections

Exercise

1. Compute the diffusivity of Ferrite for 40 equally spaced temperature from 25 ° C to 1200 ° C
2. Display results on plots of diffusivity against inverse temperature $1/T$ with different scale types

Using array operations

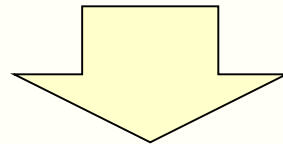
- Given the temperature vector T , how can we use the array operation to calculate

$$D = D_0 e^{-Q/RT}$$

for each value of in the vector T

- Starting point

$$\left[T(1) \quad T(2) \quad T(3) \right]$$



$$\left[D_0 e^{-\frac{Q}{RT(1)}} \quad D_0 e^{-\frac{Q}{RT(2)}} \quad D_0 e^{-\frac{Q}{RT(3)}} \right]$$

Matlab screenshots ...

1. Compute the diffusivity of Ferrite for 40 equally spaced temperature from 25 ° C to 1200 ° C

```
>>  
>> T = linspace(25,1200,40)
```

```
T =
```

```
1.0e+03 *
```

Each value in the vector is multiplied by this

```
Columns 1 through 12
```

```
0.0250 0.0551 0.0853 0.1154 0.1455 0.1756 0.2058 0.2359 0.2660 0.2962 0.3263 0.3564
```

```
Columns 13 through 24
```

```
0.3865 0.4167 0.4468 0.4769 0.5071 0.5372 0.5673 0.5974 0.6276 0.6577 0.6878 0.7179
```

```
Columns 25 through 36
```

```
0.7481 0.7782 0.8083 0.8385 0.8686 0.8987 0.9288 0.9590 0.9891 1.0192 1.0494 1.0795
```

```
Columns 37 through 40
```

```
1.1096 1.1397 1.1699 1.2000
```

Matlab screenshots ...

```
>> Do = 6.2* 10^7
```

```
Do =
```

```
62000000
```

```
>> Q = 80000
```

```
Q =
```

```
80000
```

```
>> R = 8.314
```

```
R =
```

```
8.3140
```

$$D = D_0 e^{-Q/RT}$$

- R is the ideal gas constant = 8.314 J/K/mol

- diffusion coefficient D_0 and activation energy Q depend on the material

Material	D_0 (m ² /s)	Q (J/mol)
Ferrite (α Fe)	6.2×10^{-7}	80 000

Matlab screenshots ...

```
>> D = Do * exp( -(Q/R) ./ T )
```

$$D = D_0 e^{-Q/RT}$$

```
D =
```

```
1.0e+04 *
```

Each value in the vector is multiplied by this

```
Columns 1 through 12
```

```
0.0000 0.0000 0.0000 0.0000 0.0000 0.0000 0.0000 0.0000 0.0000 0.0000 0.0000 0.0000
```

```
Columns 13 through 24
```

```
0.0000 0.0000 0.0000 0.0000 0.0000 0.0001 0.0003 0.0006 0.0014 0.0027 0.0052 0.0094
```

```
Columns 25 through 36
```

```
0.0161 0.0265 0.0419 0.0643 0.0958 0.1388 0.1965 0.2721 0.3693 0.4924 0.6457 0.8340
```

```
Columns 37 through 40
```

```
1.0624 1.3361 1.6606 2.0415
```

```
>>
```

Matlab screenshots ...

```
>> inverseT = 1 ./ T
```

```
inverseT =
```

```
Columns 1 through 12
```

```
0.0400    0.0181    0.0117    0.0087    0.0069    0.0057    0.0049    0.0042    0.0038    0.0034    0.0031    0.0028
```

```
Columns 13 through 24
```

```
0.0026    0.0024    0.0022    0.0021    0.0020    0.0019    0.0018    0.0017    0.0016    0.0015    0.0015    0.0014
```

```
Columns 25 through 36
```

```
0.0013    0.0013    0.0012    0.0012    0.0012    0.0011    0.0011    0.0010    0.0010    0.0010    0.0010    0.0009
```

```
Columns 37 through 40
```

```
0.0009    0.0009    0.0009    0.0008
```

```
>>
```

```
>>
```

```
>>
```

```
>>
```

```
>> plot( inverseT , D )
```

```
>>
```


Matlab screenshots ...

- If you do not want to display values, put ";" at the end of a line (see below). All we need is the following seven lines:

```
>>  
>> T = linspace(25,1200,40);  
>> Do = 6.2* 10^7 ;  
>> Q=80000;  
>> R = 8.314;  
>> D = Do * exp( -(Q/R) ./ T );  
>> inverseT = 1 ./ T;  
>> plot(inverseT , D );  
>>  
>>
```

Matlab screenshots ...

- If you want more clarifications, first define T, Q, R, and Do as explained in the previous slide.

```
>> T = linspace(25,1200,40);  
>> Q = 80000;  
>> R = 8.314;  
>> Do = 6.2* 10^7 ;
```

- After the above, try displaying intermediate values for your clarifications. Note that you do not need to do this in order to calculate D, this is just for your clarifications.

```
>> -(Q/R) ./ T  
>>  
>> exp( -(Q/R) ./ T )  
>>  
>> Do * exp( -(Q/R) ./ T )  
>>
```